Chem 151A, Inorganic Chemistry Spring 2015 Problem Set #3

DUE: MONDAY JUNE 1, 2015, 9:30 am

Chem 151A Mail Slot in PSB Mail Room (Next to PSB 238)

- 20 marks 1. Construct an MO diagram for $[Co(CN)_6]^{3-}$. Include the proper MO labels, AO to MO connections, electron filling, Δ_{Oct} and LFSE. For each energy level, sketch one MO (i.e. the a_{1g} , one of the three t_{1u} MOs, etc.).
- 15 marks 2. Determine the free atom term symbol(s) for a d¹ metal (ignore jj coupling).
- 15 marks 3. Using the Tanabe-Sugano diagrams on page A-38, find the values (in units of cm $^{-1}$) of D_q (or Δ_{Oct}) and B for [Cr(CN) $_6$] $^{3-}$, which has UV-Vis absorption bands at 264nm (CT), 310nm (v₁) and 378nm (v₂). All transitions are spin allowed.
- 10 marks 4. For d⁵, write all possible *spin-allowed* transitions, including the ground state and excited state molecular term symbols involved and approximate E/B values for: (i) $\Delta_{\text{Oct}}/B = 15$; (ii) $\Delta_{\text{Oct}}/B = 35$.
- 10 marks 5. The frequency of the symmetrical M-O stretching vibration of the octahedral aqua ions $[M(OH_2)_6]^{2+}$ increases along the series $Ca^{2+} < Mn^{2+} < Ni^{2+}$. How does this trend relate to acidity?
- 10 marks 6. Mn(VI) is not stable and undergoes redox. Name this phenomenon and sketch a suitable Frost diagram. Specify whether E° and ΔG° for the redox are positive, zero or negative.
- 10 marks 7. Give at least four reasons why square planar substitution reactions are associative.
- 10 marks 8. Account for the color, or lack thereof, for the following pairs of molecules:
 - (a) Cu(NH₃)₄⁺ is colorless, Cu(NH₃)₄²⁺ is intense blue
 - (b) $Co(H_2O)_6^{2+}$ is pale pink, $CoCl_4^{2-}$ is deep blue
 - (c) Au(CN)₄ and Co(CN)₆³ both form colorless crystals

Total: 100 marks